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APPENDICES

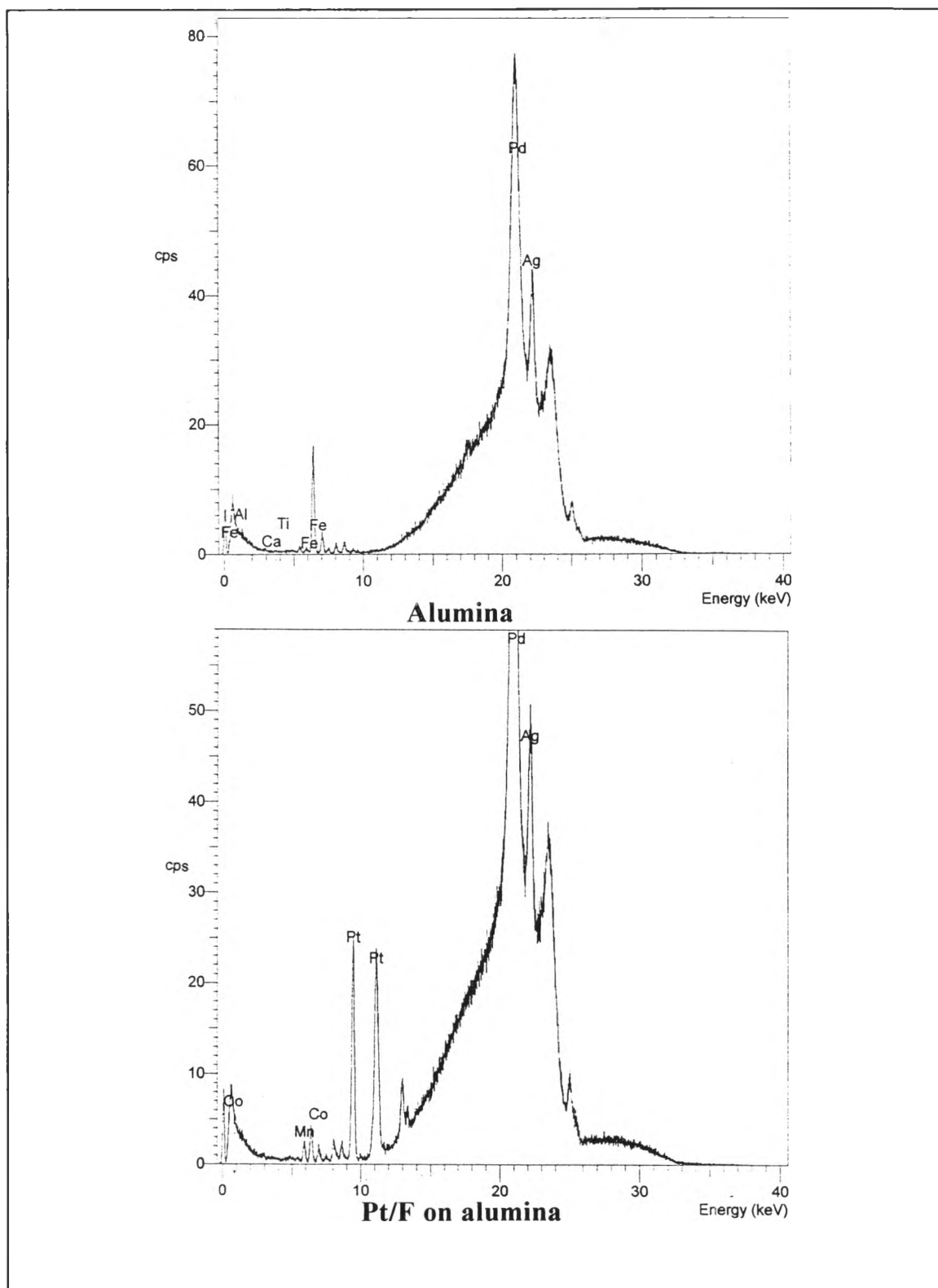


Figure A1 A plot of X-ray fluorescence data of alumina and catalyst type

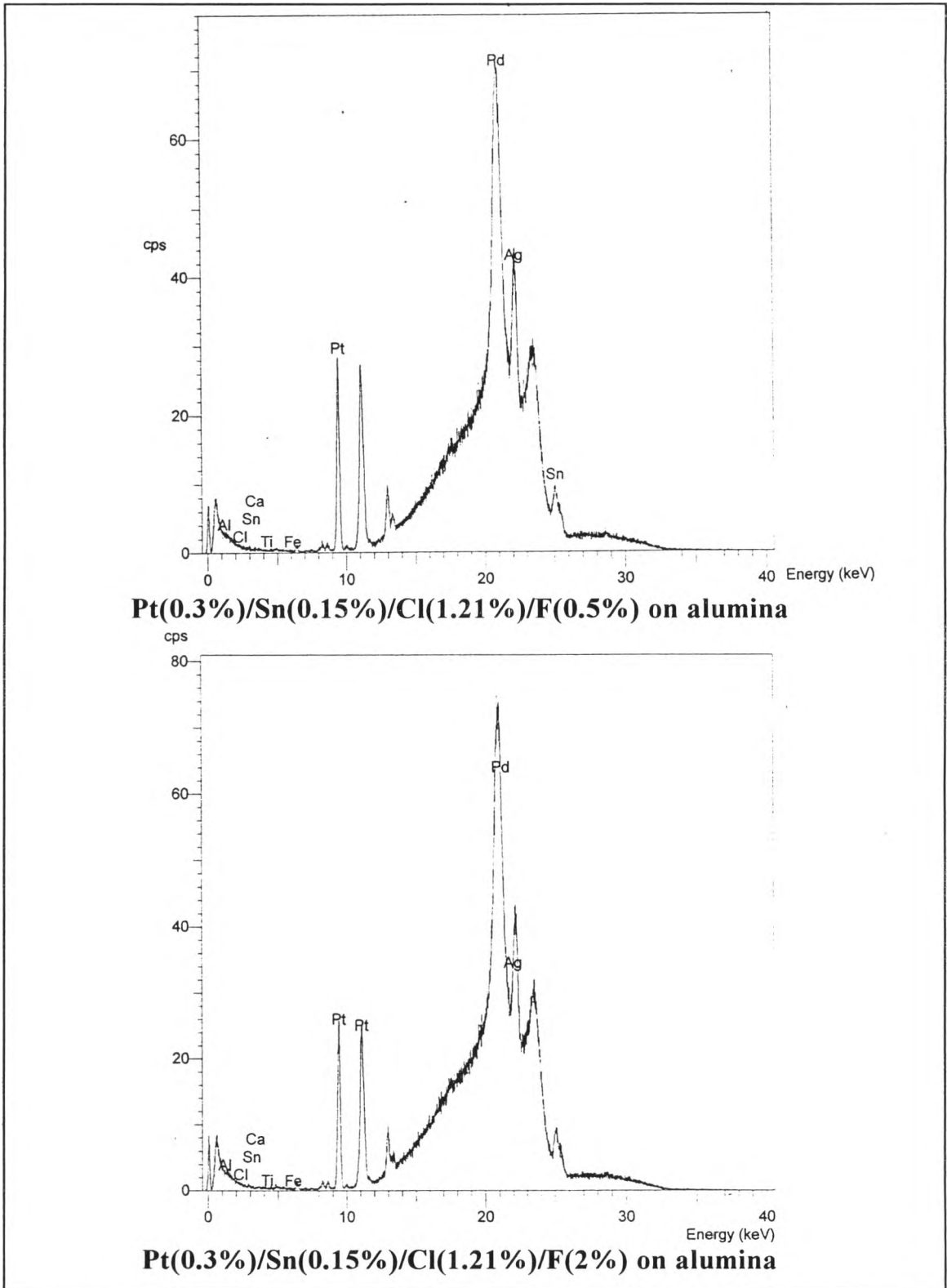


Figure A2 A plot of X-ray fluorescence data of alumina and catalyst type(continued)

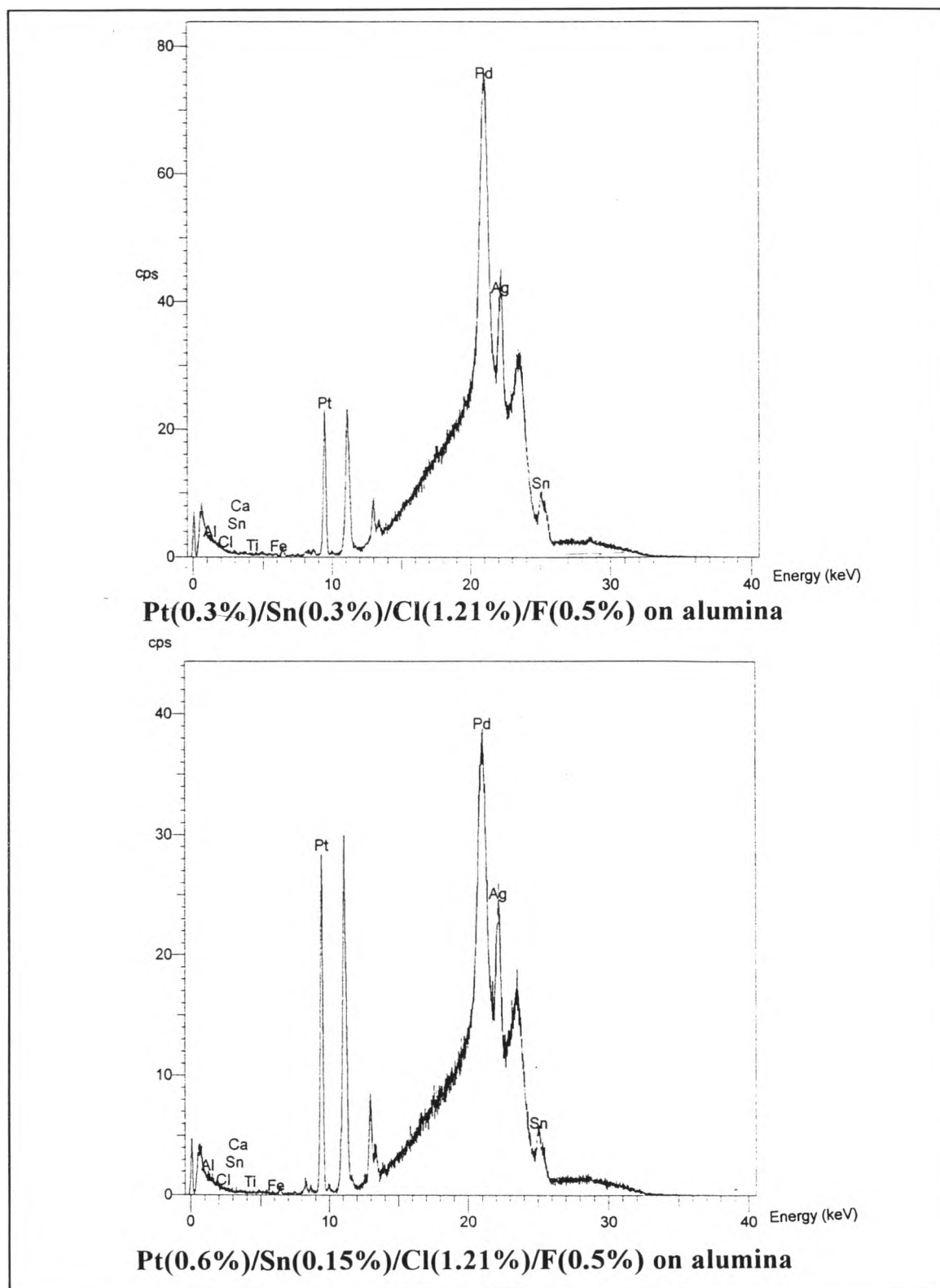


Figure A3 A plot of X-ray fluorescence data of alumina and catalyst type(continued)

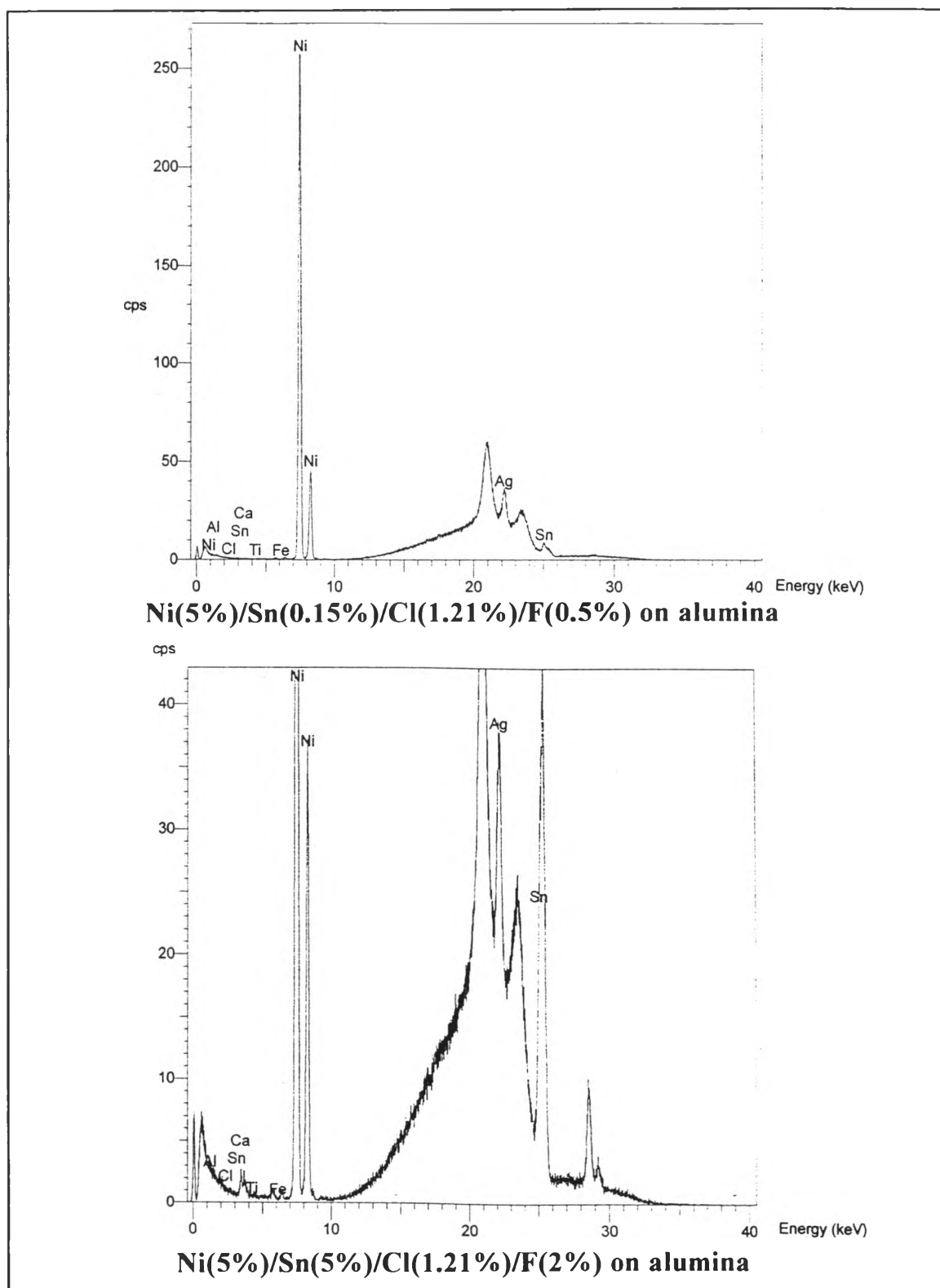


Figure A4 A plot of X-ray fluorescence data of alumina and catalyst type(continued)

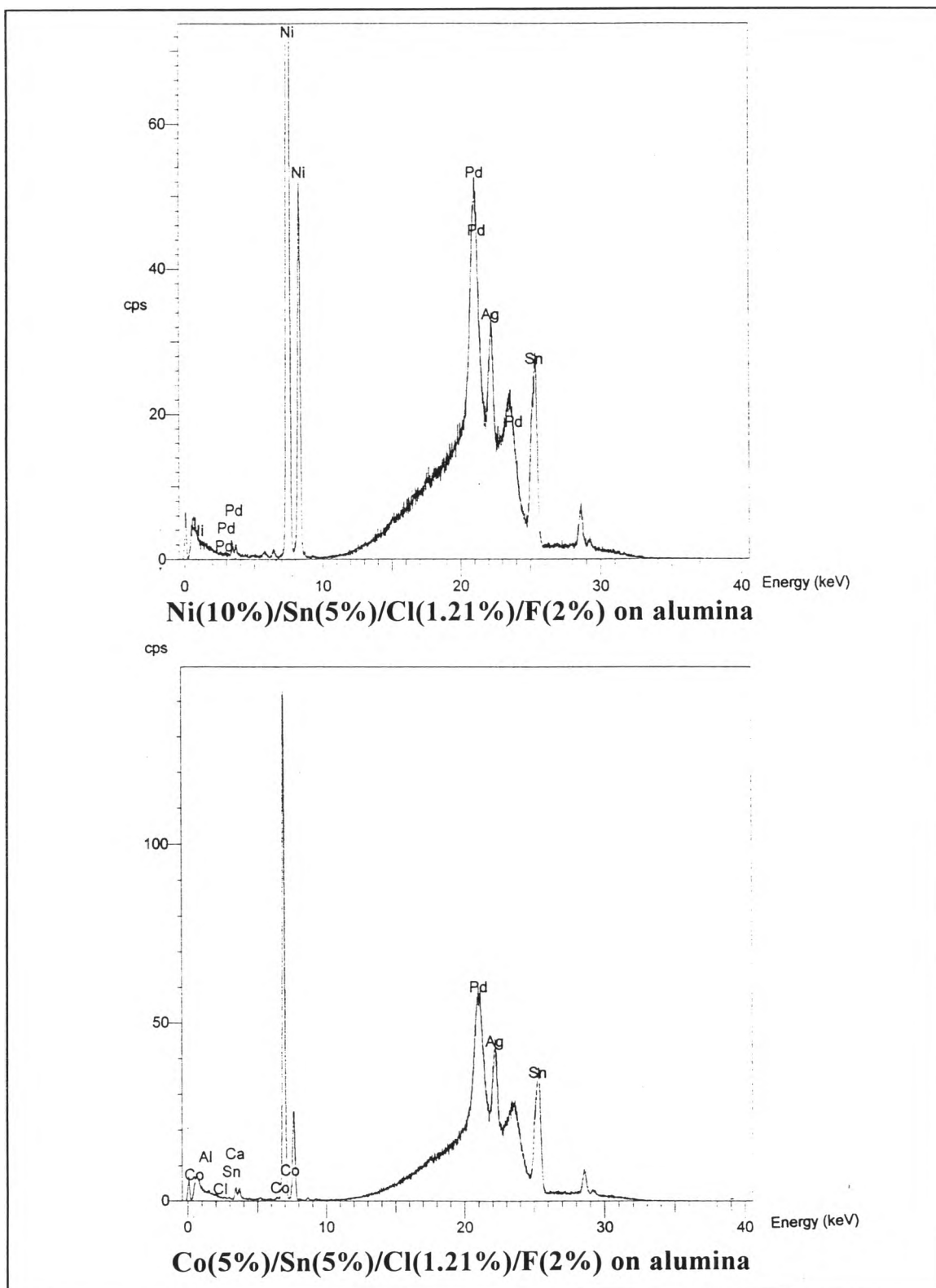


Figure A5 A plot of X-ray fluorescence data of alumina and catalyst type(continued)

Table B1 Composition of products from hydrocracking as a function of catalyst type

Molecular weight	No. of Carbon	Retention time (min)					
		Type 1	Type 3	Type 4	Type 5	Type 6	Type 7
142	isomerC ₁₀	4.37	4.37				
142	C ₁₀	4.68	4.68	4.20	4.64	4.21	4.58
156	isomerC ₁₁	5.42	5.58	4.90	5.24	4.92	5.19
156	C ₁₁	5.92	6.16	5.17	5.49	5.21	5.45
170	isomerC ₁₂	6.90	7.38	6.04	6.22	6.10	6.26
170	C ₁₂	7.49	7.93	6.59	6.78	6.65	6.76
184	isomerC ₁₃	8.58	9.14	7.78	7.84	7.87	7.92
184	C ₁₃	9.15	9.69	8.32	8.41	8.38	8.40
198	isomerC ₁₄	10.19	10.84	9.54	9.65	9.63	9.63
198	C ₁₄	10.78	11.35	10.09	10.15	10.14	10.14
212	C ₁₅	12.33	12.96	11.77	11.83	11.82	11.81
226	C ₁₆	14.02	14.75	13.44	13.54	13.51	13.52
240	C ₁₇	16.01	16.77	15.35	15.46	15.41	15.43
254	C ₁₈	18.02	18.74	17.36	17.46	17.42	17.44
268	C ₁₉	19.92	20.62	19.33	19.43	19.39	19.42
282	C ₂₀	21.67	22.42	21.20	21.29	21.28	21.28
296	C ₂₁	23.52		23.08	23.21	23.18	23.20
310	C ₂₂	25.59		25.14	25.29	25.27	25.28

Table B2 The molecular weight distributions of products from hydrocracking as a function of catalyst type

Molecular weight	No. of Carbon	% Peak Area					
		Type 1	Type 3	Type 4	Type 5	Type 6	Type 7
142	C ₁₀	20.83	14.02	4.63	9.50	4.98	8.81
156	C ₁₁	17.82	14.51	3.94	19.08	9.12	15.62
170	C ₁₂	15.01	14.81	8.57	18.07	16.50	19.62
184	C ₁₃	12.57	15.41	11.63	14.59	13.76	15.26
198	C ₁₄	9.19	12.23	13.20	8.99	11.61	10.54
212	C ₁₅	6.19	9.84	12.41	7.12	10.28	6.81
226	C ₁₆	4.50	6.46	10.54	5.00	7.88	5.45
240	C ₁₇	3.75	4.67	8.77	4.33	6.05	3.81
254	C ₁₈	3.38	3.78	7.88	4.07	5.31	3.54
268	C ₁₉	2.81	2.68	6.80	3.56	4.89	3.36
282	C ₂₀	2.06	1.59	5.71	2.71	4.06	3.09
296	C ₂₁	1.13		3.55	1.70	2.82	2.36
310	C ₂₂	0.75		2.36	1.27	2.74	1.73

Table B3 Composition of products from hydrocracking as a function of composition of Pt/Sn/Cl/F on Al₂O₃ catalyst

Molecular weight	No. of Carbon	Retention time (min)			
		Component A	Component B	Component C	Component D
142	isomerC ₁₀	4.33	4.26	4.44	4.65
142	C ₁₀	4.69	4.63	4.75	5.06
156	isomerC ₁₁	5.80	5.52	5.74	6.10
156	C ₁₁	6.13	6.07	6.22	6.45
170	isomerC ₁₂	7.40	7.33	7.42	7.85
170	C ₁₂	7.87	7.80	7.98	8.43
184	isomerC ₁₃	8.90	8.88	9.20	9.47
184	C ₁₃	9.63	9.56	9.74	10.14
198	isomerC ₁₄	10.81	10.77	10.91	11.49
198	C ₁₄	11.30	11.23	11.42	11.81
212	C ₁₅	12.92	12.85	13.04	13.50
226	C ₁₆	14.83	14.65	14.85	15.42
240	C ₁₇	16.84	16.68	16.87	17.43
254	C ₁₈	18.83	18.68	18.75	19.41
268	C ₁₉	20.75	20.56	20.75	21.28
282	C ₂₀	23.20	22.45	22.45	23.19
296	C ₂₁	25.28	24.43	24.43	25.28
310	C ₂₂	27.40	26.62	26.65	
324	C ₂₃	29.50	28.45	28.60	

Table B4 The molecular weight distributions of products from hydrocracking as a function of composition of Pt/Sn/Cl/F on Al₂O₃ catalyst

Molecular weight	No. of Carbon	% Peak Area			
		Component A	Component B	Component C	Component D
142	C ₁₀	18.00	11.22	18.57	29.61
156	C ₁₁	13.02	8.85	14.20	23.89
170	C ₁₂	12.04	10.04	13.35	15.59
184	C ₁₃	11.8	11.50	13.23	9.16
198	C ₁₄	11.07	11.41	11.53	5.58
212	C ₁₅	8.88	10.22	7.89	4.01
226	C ₁₆	6.20	7.94	5.58	3.29
240	C ₁₇	5.11	7.21	4.61	2.72
254	C ₁₈	4.38	6.39	3.64	2.43
268	C ₁₉	3.53	5.47	2.79	1.86
282	C ₂₀	2.19	4.11	1.82	1.14
296	C ₂₁	1.58	2.65	1.21	0.72
310	C ₂₂	1.22	1.82	0.97	
324	C ₂₃	0.97	1.19	0.61	

Table B5 Composition of products from hydrocracking as a function of catalyst concentration

Molecular weight	No. of Carbon	Retention time (min)		
		30%	35%	40%
142	isomerC ₁₀	4.94	4.65	5.28
142	C ₁₀	5.22	5.06	5.51
156	isomerC ₁₁	6.28	6.12	6.19
156	C ₁₁	6.6	6.45	6.65
170	isomerC ₁₂	7.8	7.85	7.76
170	C ₁₂	8.29	8.43	8.27
184	isomerC ₁₃	9.52	9.47	9.60
184	C ₁₃	10.03	10.14	10.17
198	isomerC ₁₄	11.21	11.49	11.58
198	C ₁₄	11.69	11.81	12.15
212	C ₁₅	13.36	13.50	14.07
226	C ₁₆	15.25	15.42	16.18
240	C ₁₇	17.37	17.43	
254	C ₁₈	19.33	19.41	
268	C ₁₉	21.20	21.28	
282	C ₂₀	23.08	23.19	
296	C ₂₁	25.13	25.28	

Table B6 The molecular weight distributions of products from hydrocracking as a function of catalyst concentration

Molecular weight	No. of Carbon	% Peak Area		
		30%	35%	40%
142	C ₁₀	17.58	29.61	31.53
156	C ₁₁	23.58	23.89	28.65
170	C ₁₂	16.72	15.59	17.30
184	C ₁₃	12.65	9.16	10.99
198	C ₁₄	7.29	5.58	5.77
212	C ₁₅	5.89	4.01	3.60
226	C ₁₆	4.18	3.29	2.16
240	C ₁₇	3.43	2.72	
254	C ₁₈	3.11	2.43	
268	C ₁₉	2.47	1.86	
282	C ₂₀	1.93	1.14	
296	C ₂₁	1.18	0.72	

Table B7 Composition of products from hydrocracking as a function of reaction time

Molecular weight	No. of Carbon	Retention time (min)			
		4 hours	8 hours	10 hours	12 hours
142	isomerC ₁₀		4.49	4.49	5.28
142	C ₁₀	4.72	4.78	4.92	5.51
156	isomerC ₁₁	5.66	5.75	5.79	6.19
156	C ₁₁	6.05	6.19	6.28	6.65
170	isomerC ₁₂	7.21	7.29	7.46	7.76
170	C ₁₂	7.71	7.91	8.03	8.27
184	isomerC ₁₃	8.92	9.15	9.23	9.60
184	C ₁₃	9.44	9.68	9.79	10.17
198	isomerC ₁₄	10.60	10.97	10.92	11.58
198	C ₁₄	11.09	11.35	11.45	12.15
212	C ₁₅	12.69	12.97	13.07	14.07
226	C ₁₆	14.45	14.79	14.89	16.18
240	C ₁₇	16.47	16.53	16.82	
254	C ₁₈	18.47	18.80	18.82	
268	C ₁₉	20.35	20.67	20.73	
282	C ₂₀	22.14	22.48	22.62	
296	C ₂₁	24.08			
310	C ₂₂	26.20			
324	C ₂₃	28.28			
338	C ₂₄	30.32			

Table B8 The molecular weight distributions of products from hydrocracking as a function of reaction time

Molecular weight	No. of Carbon	% Peak Area			
		4 hours	8 hours	10 hours	12 hours
142	C ₁₀	0.81	17.98	21.48	17.48
156	C ₁₁	4.05	16.40	17.48	15.88
170	C ₁₂	9.72	12.65	13.49	9.59
184	C ₁₃	12.96	10.47	10.99	6.09
198	C ₁₄	12.96	10.28	9.09	3.20
212	C ₁₅	11.74	8.50	7.79	2.00
226	C ₁₆	9.85	6.52	5.49	1.20
240	C ₁₇	7.83	5.53	5.19	
254	C ₁₈	7.56	4.74	3.70	
268	C ₁₉	6.34	3.95	3.10	
282	C ₂₀	5.40	2.96	2.20	
296	C ₂₁	3.91			
310	C ₂₂	3.10			
324	C ₂₃	2.56			
338	C ₂₄	1.21			

Table B9 Composition of products from hydrocracking as a function of reaction temperature

Molecular weight	No. of Carbon	Retention time (min)	
		350°C	400°C
142	isomerC ₁₀		5.28
142	C ₁₀	5.12	5.51
156	isomerC ₁₁	6.09	6.19
156	C ₁₁	6.55	6.65
170	isomerC ₁₂	7.74	7.76
170	C ₁₂	8.26	8.27
184	isomerC ₁₃	9.49	9.60
184	C ₁₃	10.04	10.17
198	isomerC ₁₄	11.18	11.58
198	C ₁₄	11.71	12.15
212	C ₁₅	13.38	14.07
226	C ₁₆	15.27	16.18
240	C ₁₇	17.28	
254	C ₁₈	19.26	
268	C ₁₉	21.13	
282	C ₂₀	23.08	
296	C ₂₁	25.10	

Table B10 The molecular weight distributions of products from hydrocracking as a function of reaction temperature

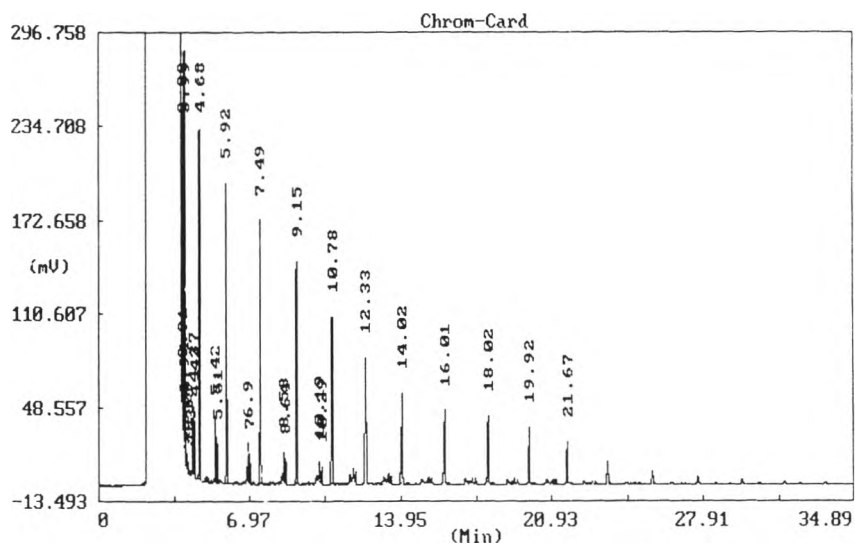
Molecular weight	No. of Carbon	% Peak Area	
		350°C	400°C
142	C ₁₀	3.81	31.53
156	C ₁₁	8.71	28.65
170	C ₁₂	12.75	17.30
184	C ₁₃	13.73	10.99
198	C ₁₄	10.89	5.77
212	C ₁₅	8.82	3.60
226	C ₁₆	7.84	2.16
240	C ₁₇	6.75	
254	C ₁₈	5.56	
268	C ₁₉	4.14	
282	C ₂₀	2.51	
296	C ₂₁	1.20	

Table B11 Composition of products from hydrocracking as a function of hydrogen pressure

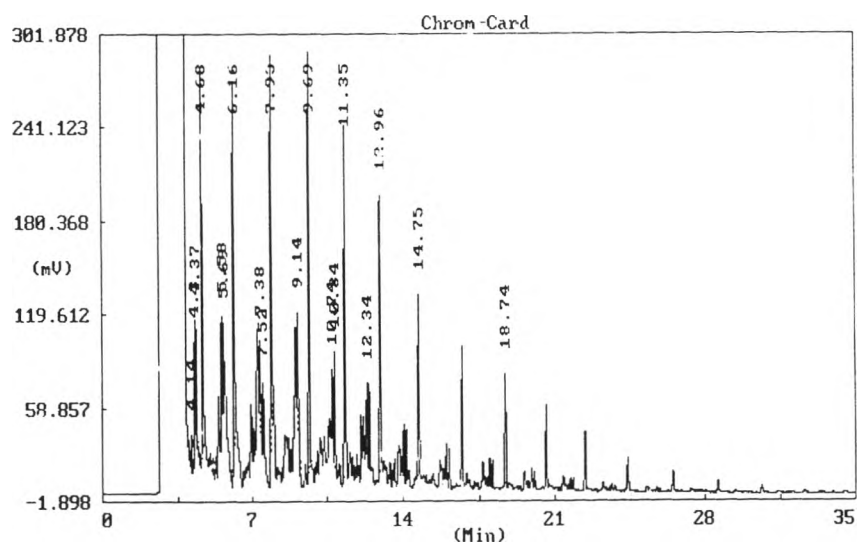
Molecular weight	No. of Carbon	Retention time (min)	
		500 psig	600 psig
142	isomerC ₁₀	4.80	5.28
142	C ₁₀	5.24	5.51
156	isomerC ₁₁	6.05	6.19
156	C ₁₁	6.51	6.65
170	isomerC ₁₂	7.71	7.76
170	C ₁₂	8.27	8.27
184	isomerC ₁₃	9.47	9.60
184	C ₁₃	10.05	10.17
198	isomerC ₁₄	11.18	11.58
198	C ₁₄	11.72	12.15
212	C ₁₅	13.39	14.07
226	C ₁₆	15.28	16.18
240	C ₁₇	11.29	
254	C ₁₈	19.26	
268	C ₁₉	21.13	
282	C ₂₀	23.02	
296	C ₂₁	25.13	

Table B12 The molecular weight distributions of products from hydrocracking as a function of hydrogen pressure

Molecular weight	No. of Carbon	% Peak Area	
		500 psig	600 psig
142	C ₁₀	9.21	31.53
156	C ₁₁	8.31	28.65
170	C ₁₂	11.69	17.30
184	C ₁₃	13.48	10.99
198	C ₁₄	12.58	5.77
212	C ₁₅	10.56	3.60
226	C ₁₆	8.54	2.16
240	C ₁₇	7.87	
254	C ₁₈	6.52	
268	C ₁₉	5.39	
282	C ₂₀	3.82	
296	C ₂₁	2.02	

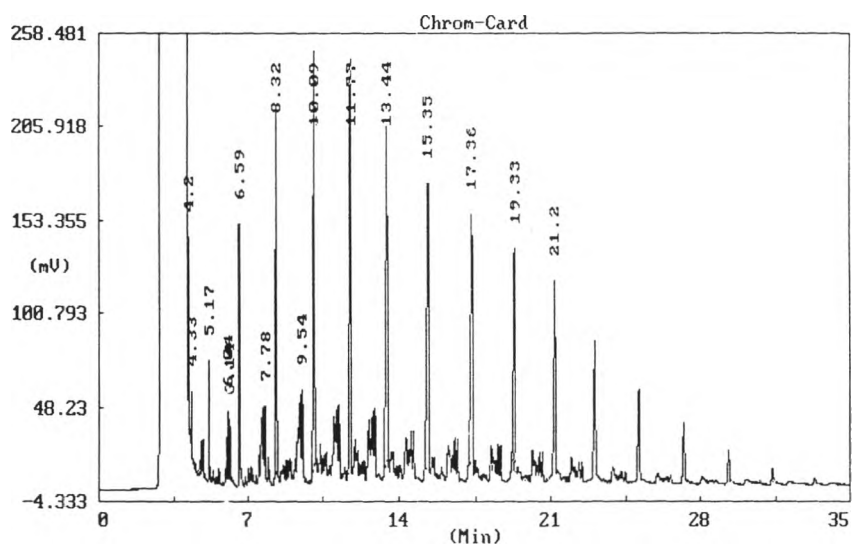


Type 1: Commercial catalyst

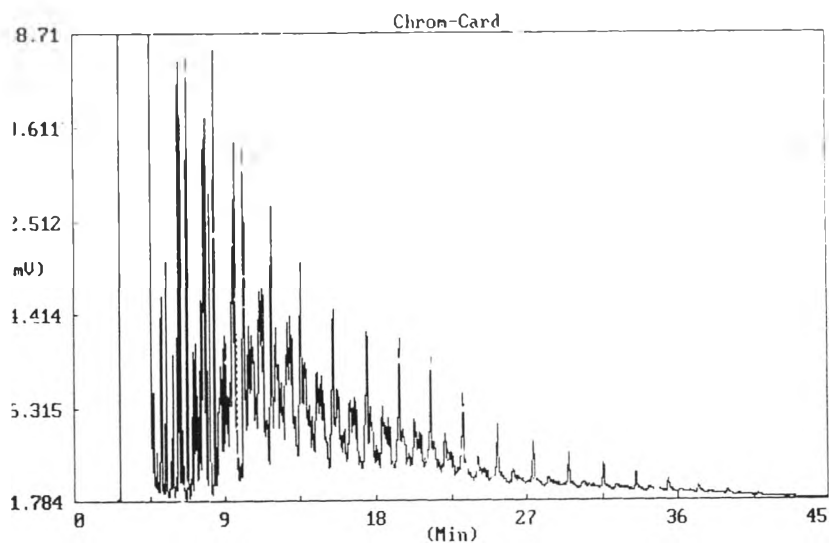


Type 3: Pt(0.3%)/Sn(0.15%)/Cl(1.21%)/F(0.5%) on Al₂O₃

Figure B1 GC/MS chromatogram of products from hydrocracking as a function of catalyst type

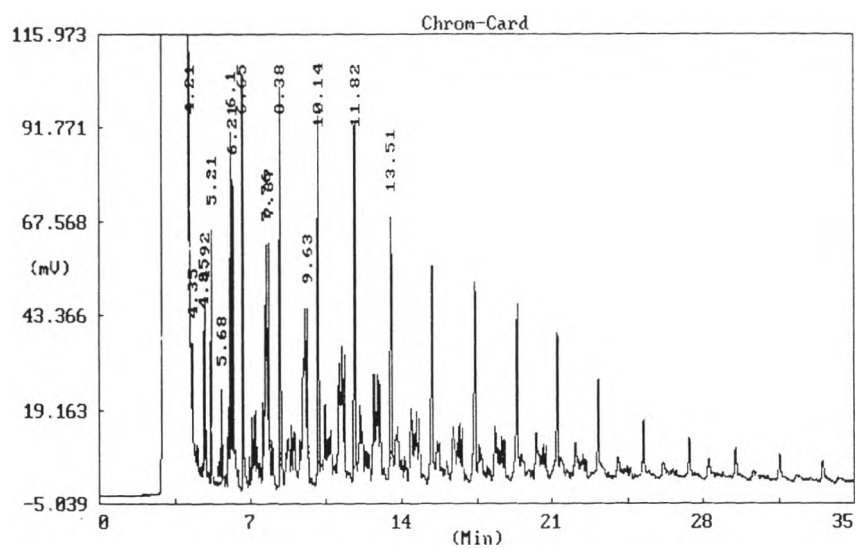


Type 4: Ni(5%)/Sn(0.15%)/Cl(1.21%)/F(0.5%) on Al₂O₃

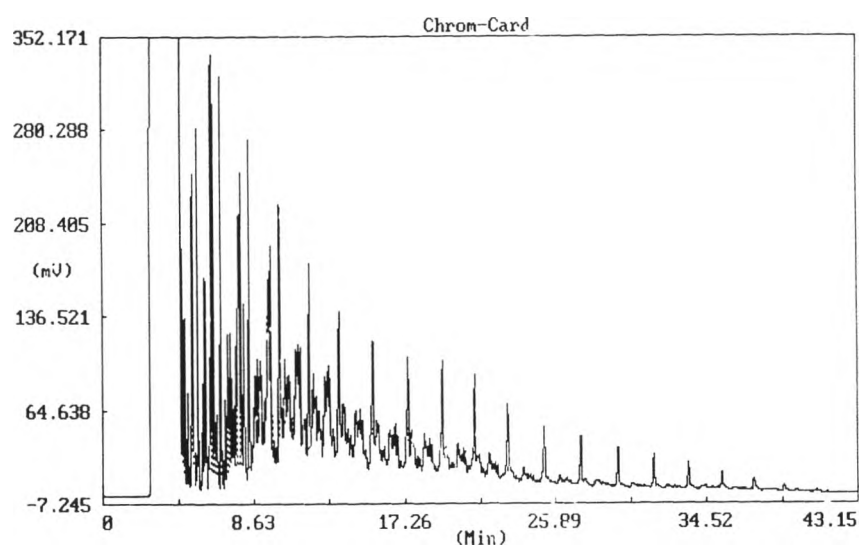


Type 5: Ni(5%)/Sn(5%)/Cl(1.21%)/F(2%) on Al₂O₃

Figure B2 GC/MS chromatogram of products from hydrocracking as a function of catalyst type (continued)

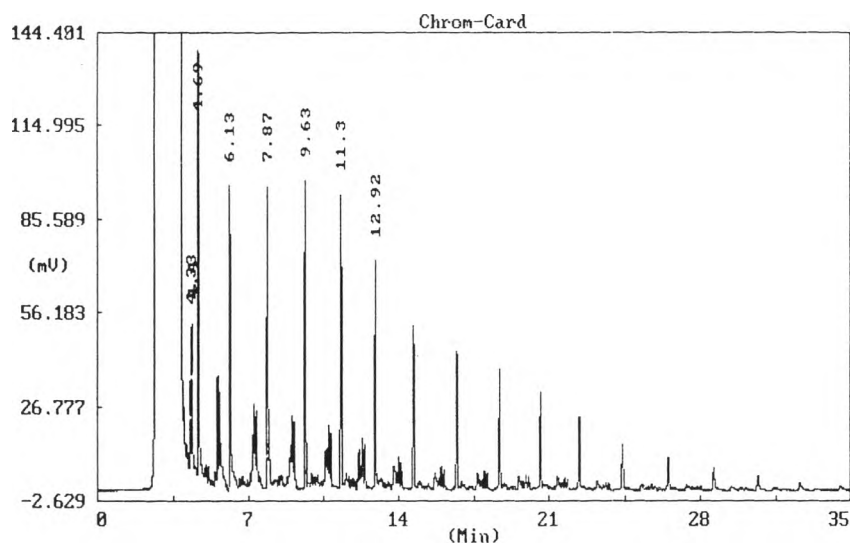


Type 6: Ni(10%)/Sn(5%)/Cl(1.21%)/F(2%) on Al₂O₃

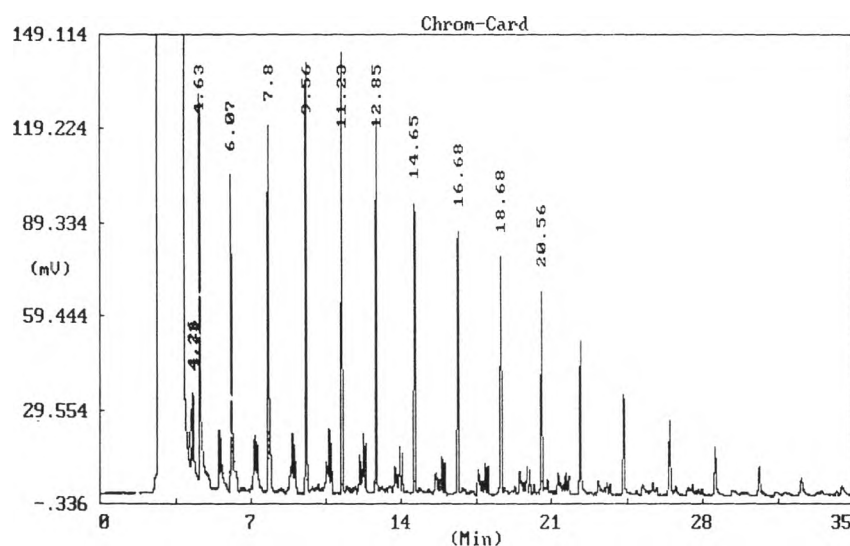


Type 7: Co(5%)/Sn(5%)/Cl(1.21%)/F(2%) on Al₂O₃

Figure B3 GC/MS chromatogram of products from hydrocracking as a function of catalyst type (continued)

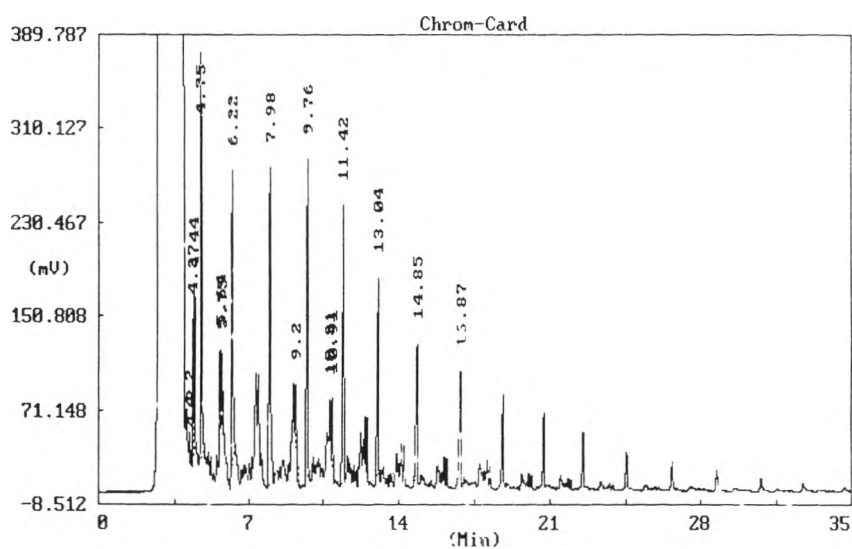


Component A: Pt(0.3%)/Sn(0.15%)/Cl(1.21%)/F(0.5%) on Al₂O₃

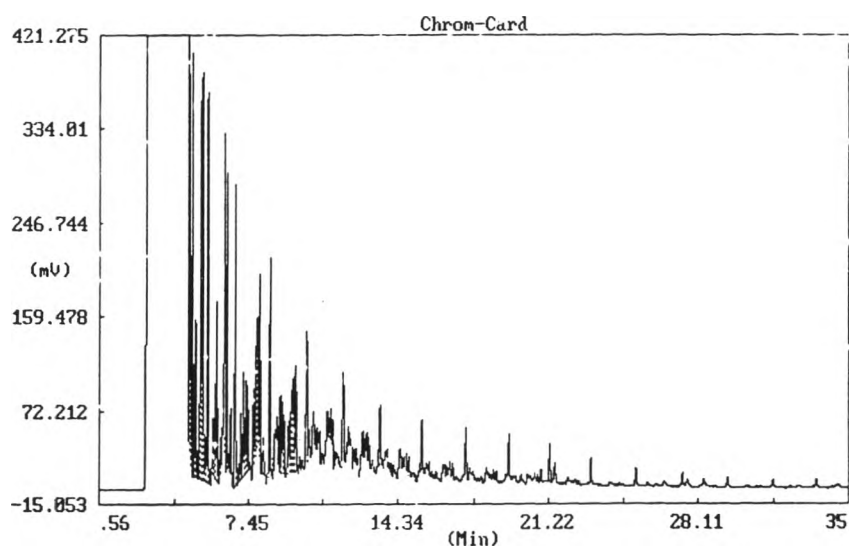


Component B: Pt(0.3%)/Sn(0.15%)/Cl(1.21%)/F(2%) on Al₂O₃

Figure B4 GC/MS chromatogram of products from hydrocracking as a function of %component of Pt/Sn/Cl/F on Al₂O₃ catalyst



Component C: Pt(0.3%)/Sn(0.3%)/Cl(1.21%)/F(0.5%) on Al₂O₃



Component D: Pt(0.6%)/Sn(0.15%)/Cl(1.21%)/F(0.5%) on Al₂O₃

Figure B5 GC/MS chromatogram of products from hydrocracking as a function of %component of Pt/Sn/Cl/F on Al₂O₃ catalyst (continued)

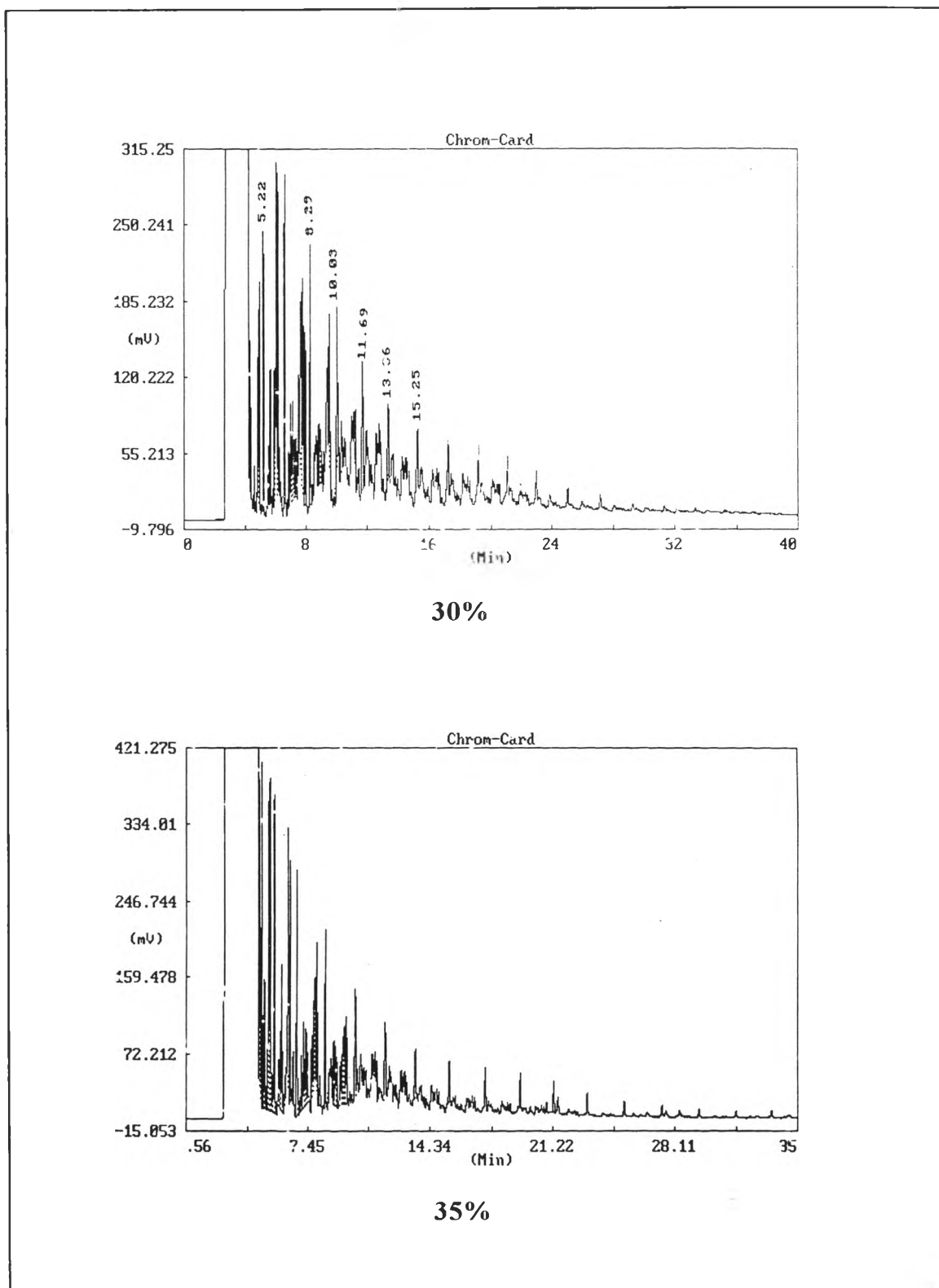


Figure B6 GC/MS chromatogram of products from hydrocracking as a function of catalyst concentration

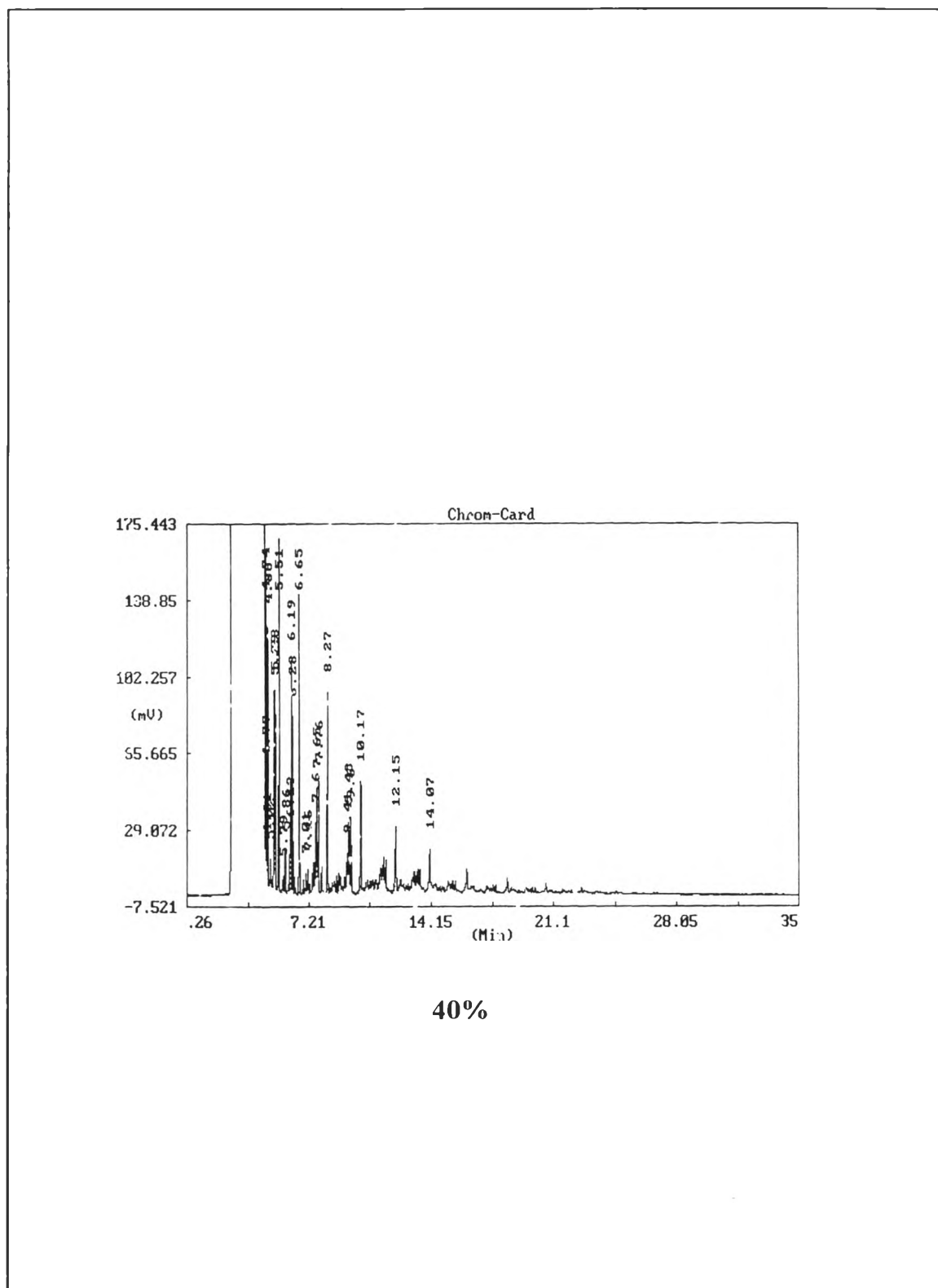


Figure B7 GC/MS chromatogram of products from hydrocracking as a function of catalyst concentration (continued)

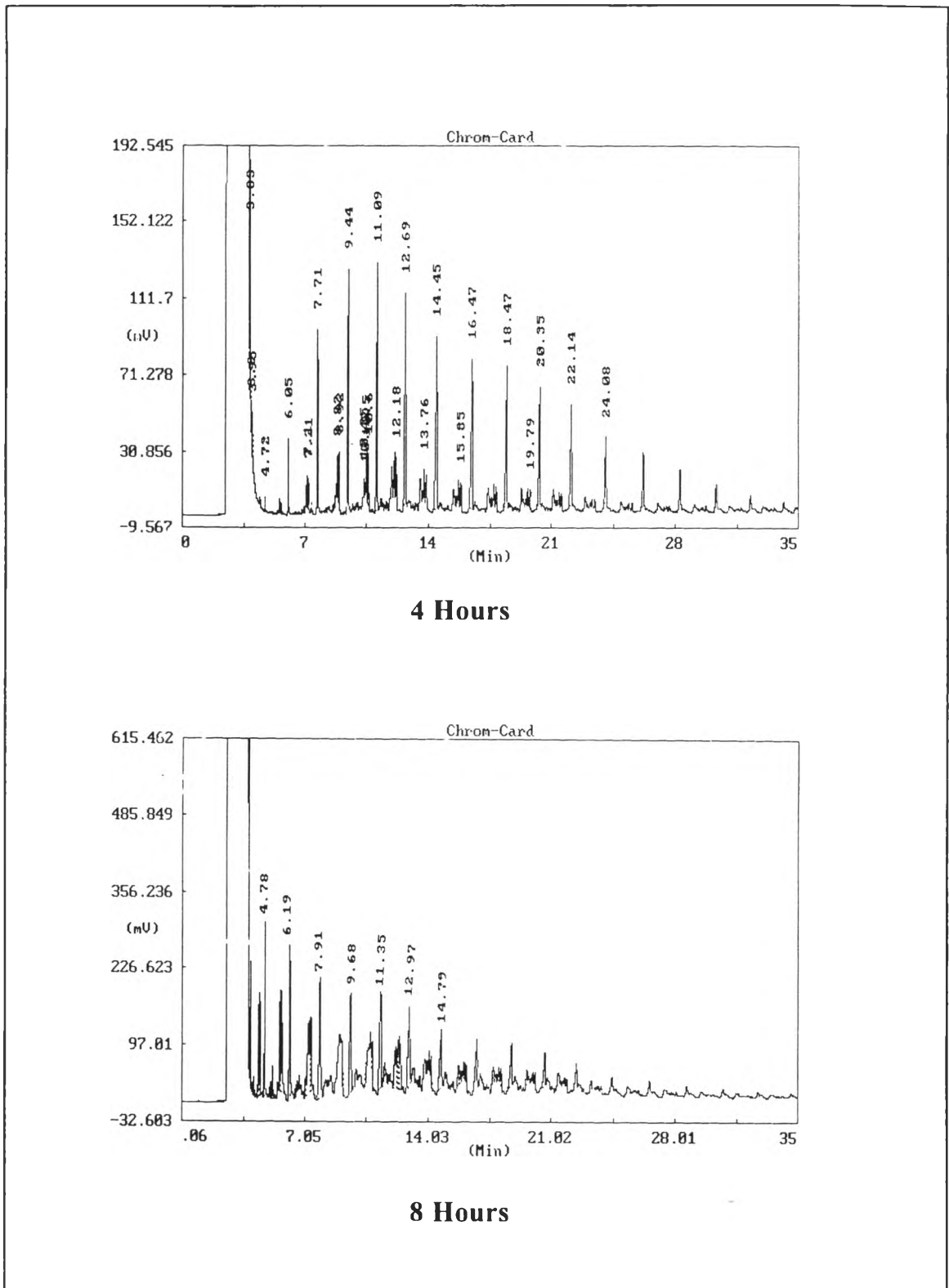


Figure B8 GC/MS chromatogram of products from hydrocracking as a function of reaction time

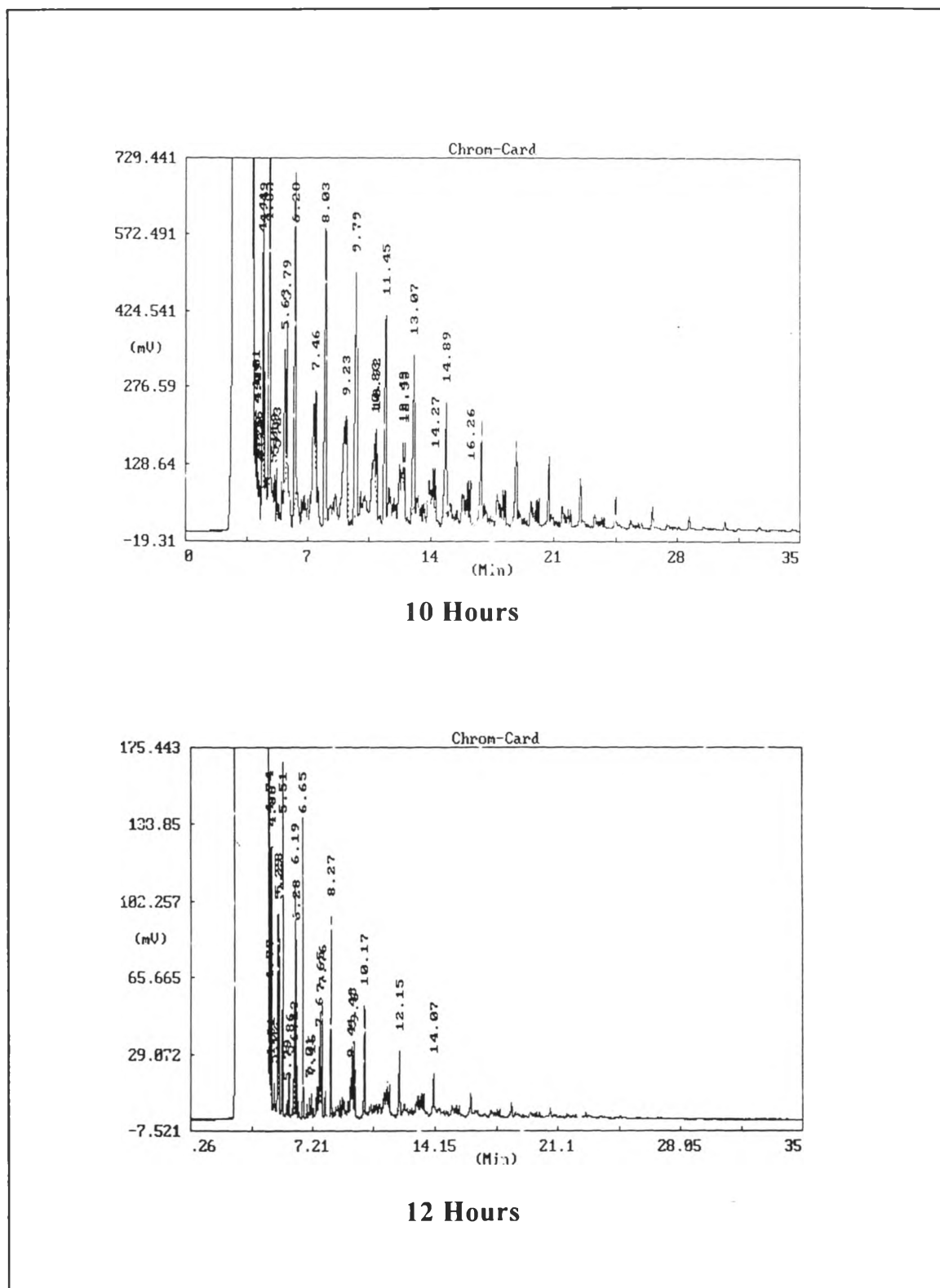


Figure B9 GC/MS chromatogram of products from hydrocracking as a function of reaction time (continued)

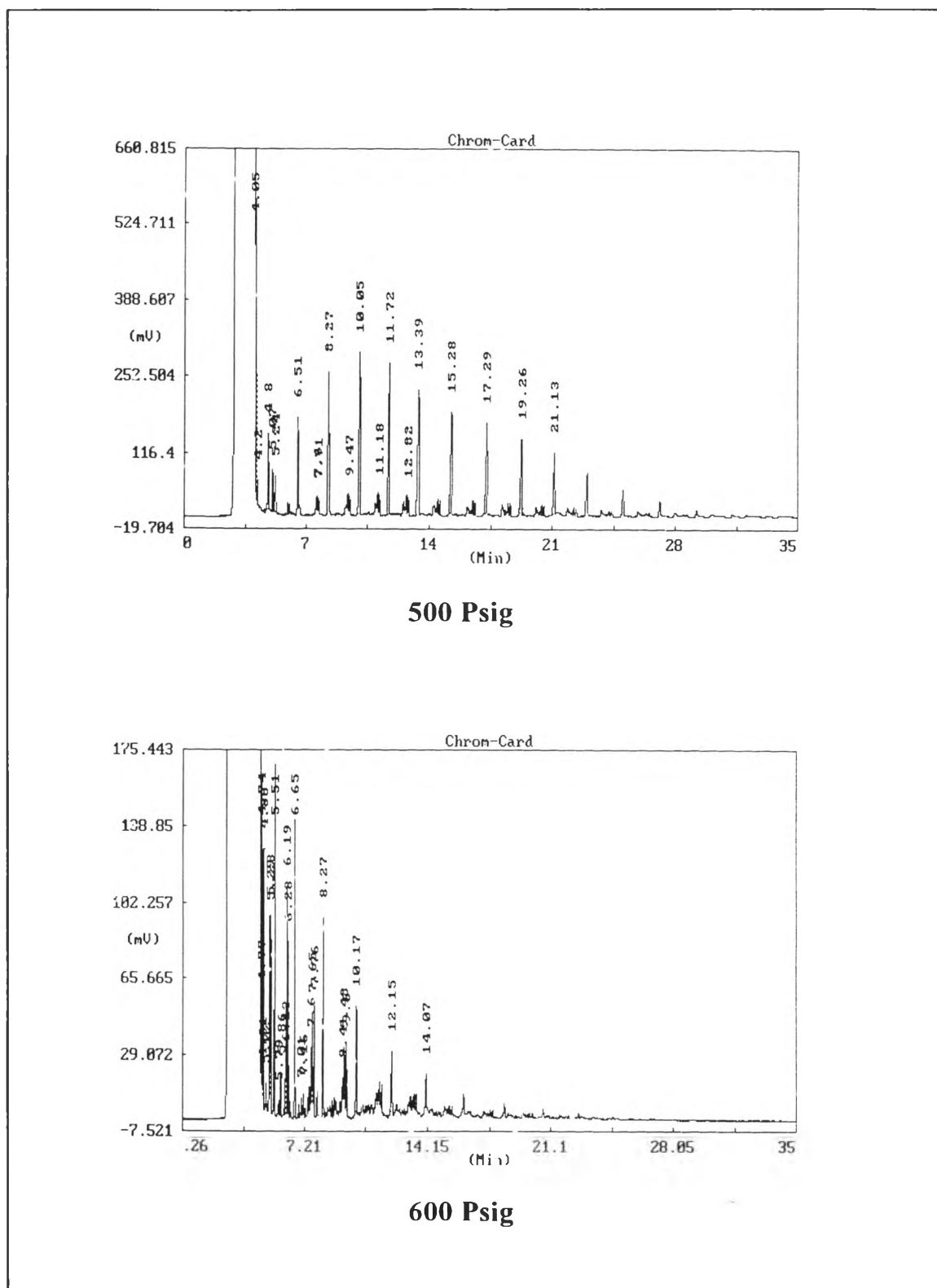


Figure B11 GC/MS chromatogram of products from hydrocracking as a function of hydrogen pressure

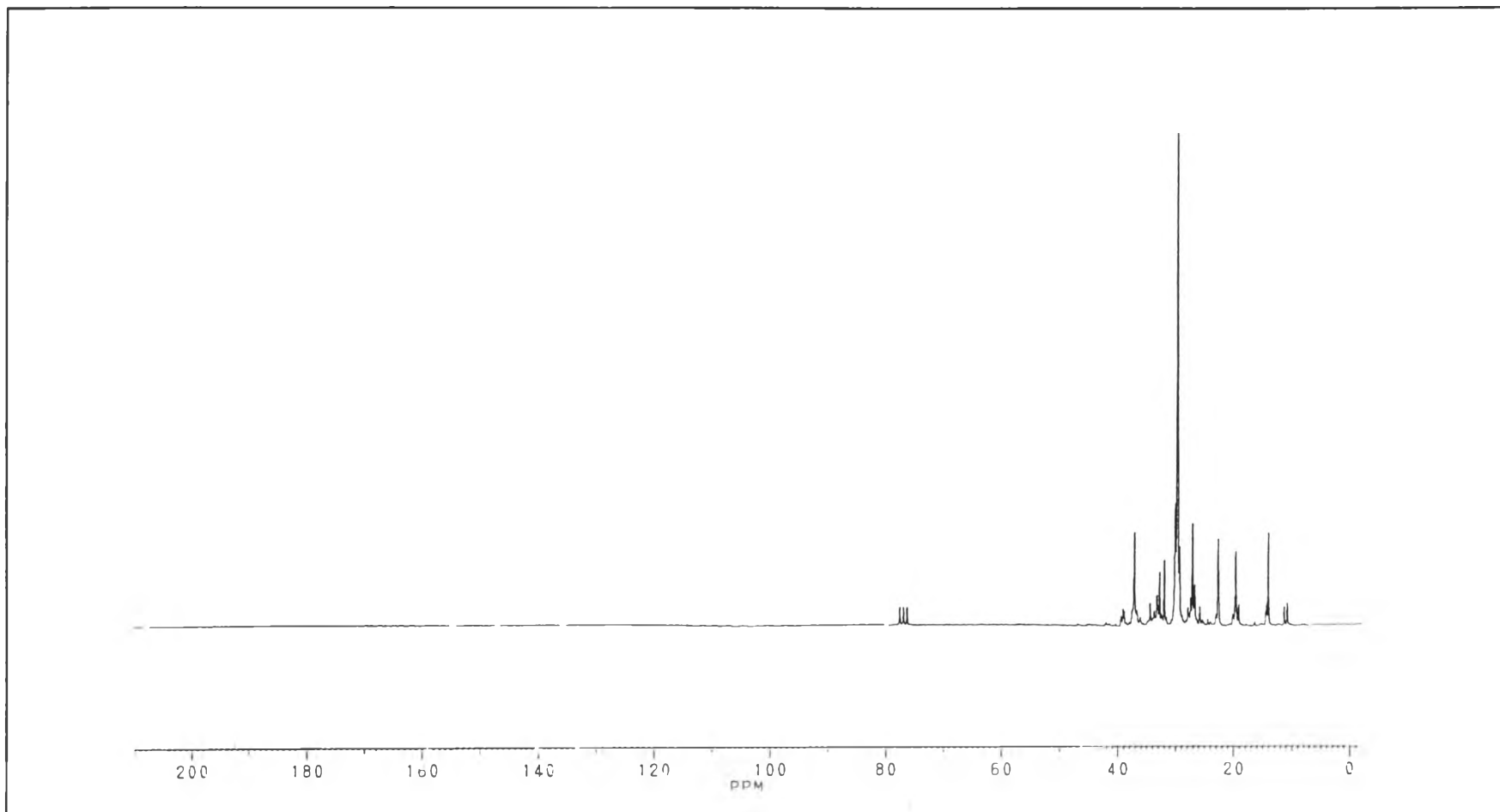


Figure B12 ^{13}C NMR (CDCl_3) spectrum of product from hydrocracking of used polyethylene at optimum condition (40%wt. of Pt(0.6%)/Sn(0.15%)/Cl(1.21%)/F(0.5%) on Al_2O_3 catalyst, 400°C , 600 psig, 12 hours)

VITA

Miss Ruthai Leesuksan was born on September 24, 1972 in Bangkok, Thailand. She received her Bachelor of Science degree in Chemistry, Srinakarinwirot University Southern Campus in 1994. She continued her studies towards her Master's degree at Chulalongkorn University, Multidisciplinary program of Petrochemistry and Polymer, Graduate School, in 1994 and completed the program in 1997.

